

Chemistry Notes 2013 – Titration Practical Test

General

- Titration is a form of volumetric analysis – it's used to determine the concentration of an unknown solution by reacting it with another solution of known concentration.
 - o 1 solution's concentration is known
 - o Measure the volumes of both exactly, then determine by means of calculations
- Indicator to be used: Methyl orange
 - o Acid: Orange
 - o Alkali: Yellow
- Note: A fixed quantity of the unknown solution, is titrated with the standard solution, and **the volume of standard solution used is measured.**

Technique

- Wash all apparatus thoroughly with tap water, and then rinse with distilled water.
- Wash burette and pipette with respective solutions (conical flask with distilled water only).
- Read all readings from lower meniscus.
- Add 3 drops of methyl orange.
- End point for methyl orange is: (ta-da) orange.
- Prepare both conical flasks before titrating to compare colour

Recording and Answering

- Burette: 2dp, Pipette: 1dp
- Tick at least 2 titration readings (ensure they do not differ by more than 0.1cm^3)
- All calculations to 3sf.

Titration Number	1	2
Final Reading / cm^3		
Initial Reading / cm^3		
Volume of P used / cm^3		
Best titration results (v)		

Volume of Q used = _____

Average volume of P used = _____ (take average based on ticked best titration results)